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## The Crystal Structure of 5-Acetoxy-6-methoxy-8-nitroquinoline

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The structure of 5-acetoxy-6-methoxy-8-nitroquinoline was determined from three-dimensional equi-inclination Weissenberg data by a direct method. The compound crystallizes in the monoclinic system with space group  $P2_1/c$ . The cell data are:  $a = 7.923 \pm 0.003$ ,  $b = 9.056 \pm 0.003$ ,  $c = 16.694 \pm 0.006$  Å,  $\beta = 97.40 \pm 0.02^\circ$ ,  $D_x = 1.466$  g.cm<sup>-3</sup>. The bond lengths are in agreement with those calculated by the valence bond theory. A comparison of the observed with the calculated exocyclic bond lengths indicates that only a small amount of resonance interaction occurs between the quinoline ring and the substituent groups. However, the bond lengths of the ring are unaffected by the resonance interaction with the substituent groups in that the observed values agree with theoretical bond distances of the unsubstituted molecule obtained in a recent self-consistent-field molecular-orbital type calculation. The exocyclic COC bond angles average  $117^\circ$  instead of the normal tetrahedral value owing to steric repulsions between some atoms in the methoxy and acetoxy groups and some of those in the ring. Packing and intramolecular steric requirements force the nitro and acetoxy groups to rotate out of the plane of the ring by  $59^\circ$  and  $79^\circ$  respectively. The resonance energy of the acetoxy group is sufficient to maintain its planarity in the overcrowded environment. On the contrary the quinoline ring deviates slightly but significantly from planarity.

### Introduction

5-Acetoxy-6-methoxy-8-nitroquinoline (Fig. 1) is a synthetic intermediate in the preparation of 5-hydroxy-6-methoxyquinoline, a reference compound for use in a study of 6-methoxyquinoline metabolites (Sax & Lynch, 1966). Its spectral and chemical properties and those of 5-hydroxy-6-methoxy-8-nitroquinoline have been determined (Griffin, Byrne, Sax & Lynch, 1966). The differences in the properties of these compounds suggest a greater diversity in their molecular structures than is implied by the mere replacement of an acetoxy by a hydroxy group. In order to ascertain the structural dissimilarities in these two molecules, both were selected for crystal structure analysis, but only the structure of the acetoxy compound is considered in this paper.

### Crystal data

5-Acetoxy-6-methoxy-8-nitroquinoline, C<sub>12</sub>H<sub>10</sub>O<sub>5</sub>N<sub>2</sub>,  
 M.W. 262.23  
 Monoclinic, space group  $P2_1/c$   
 $a = 7.923 \pm 0.003$ ,  $b = 9.056 \pm 0.003$ ,  $c = 16.694 \pm 0.006$  Å  
 $\beta = 97.40 \pm 0.02^\circ$   
 $D_x = 1.466$  g.cm<sup>-3</sup>,  $D_m = 1.472$  g.cm<sup>-3</sup>  
 $Z = 4$

### Experimental

Crystals of the acetoxy compound were provided by Dr S. M. Sax of the Department of Pathology, Western Pennsylvania Hospital, Pittsburgh, Pa. The unit-cell dimensions were measured on a Picker full circle diffractometer with Cu  $K\alpha$  radiation (1.5418 Å). The X-ray intensity data were collected on multifilm equi-inclination Weissenberg photographs with nickel-filtered Cu  $K\alpha$  radiation from crystals rotated about the  $a$  and  $b$  axes. Both crystals were cut to approximate a

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rectangular parallelepiped in shape. The dimensions of the one that was mounted about the  $b$  axis were 0.45, 0.55 and 1.15 mm in the  $a^*$ ,  $c$ , and  $b$  directions respectively. The dimensions of the crystal rotated about the  $a$  axis were 0.3 mm in the  $b$  and  $c^*$  directions and 0.75 mm along  $a$ . Seven layers were measured on the former and six on the latter crystal. The intensity data were estimated visually by comparison with a standard scale. Interfilm and interlayer scaling of the data were performed on the IBM 1620 computer using a series of programs (University of Pittsburgh Technical Reports, 1963). The Lorentz, polarization, and spot shape corrections were applied but none were made for absorption or extinction. The 1930 observed  $F$ 's were scaled by a Wilson plot and the signs of 420

normalized structure factors were uniquely determined by a direct method procedure programmed for the IBM 1620 computer (Beurskens, 1963). An  $E$ -map synthesis gave well-resolved peaks indicating the atomic positions of the molecule. Structure factors calculated with these coordinates (assuming an overall temperature factor of  $3.0 \text{ \AA}^2$ ) agreed reasonably well at this stage with the observed data, giving a reliability index ( $R$ ) of 0.34.

The entire observed data, except for nine strong low order reflections probably affected by extinction, were included in the full-matrix least-squares refinement using the Busing, Martin & Levy (1962) program. The atomic form factors were taken from *International Tables for X-ray Crystallography* (1962). The Hughes (1941) weighting scheme with  $4F_{\min}$  equal to 6.0 was

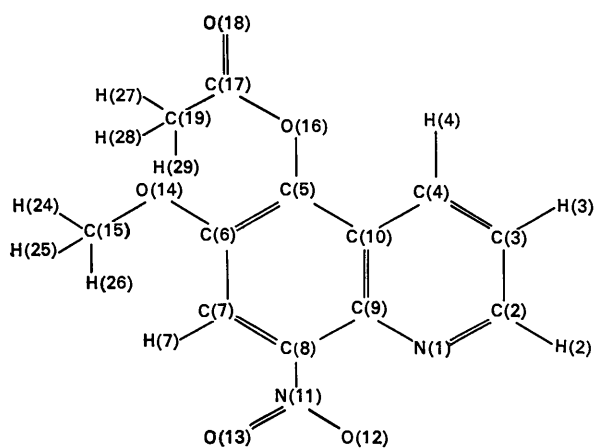


Fig. 1. Atomic numbering in 5-acetoxy-6-methoxy-8-nitroquinoline.

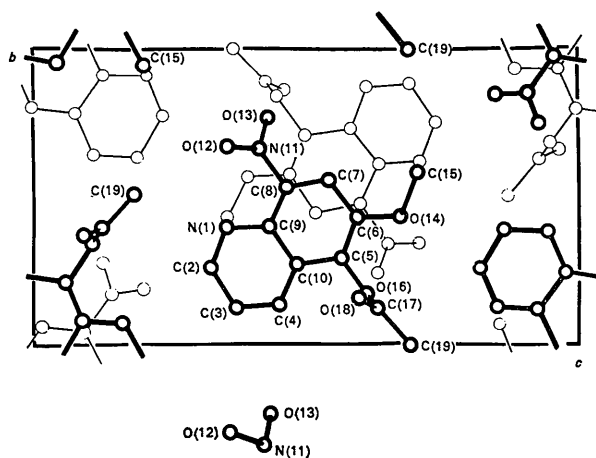


Fig. 2. Molecular packing viewed along normal to the  $bc$  plane.

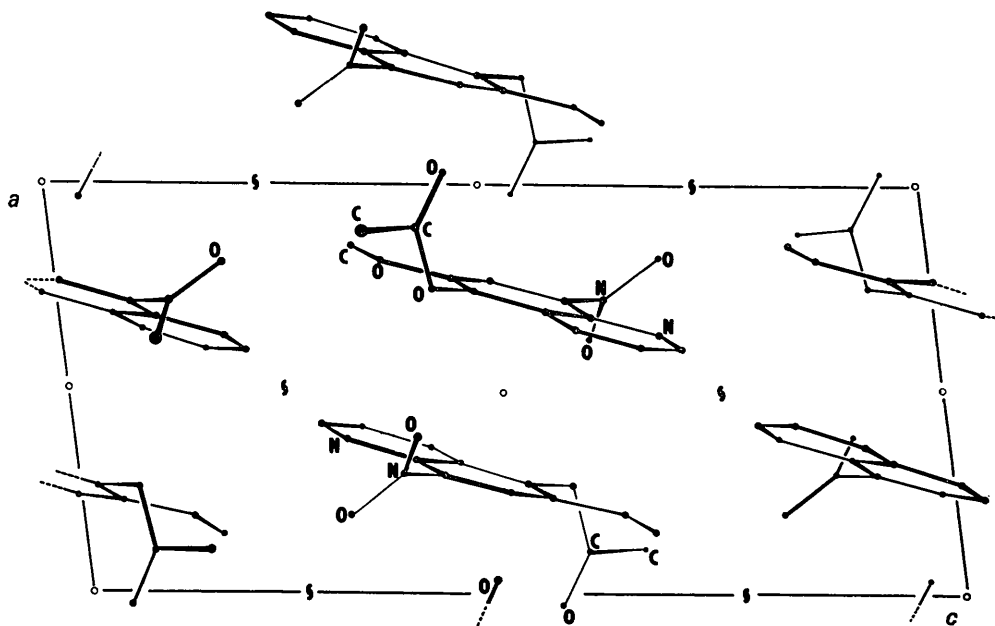


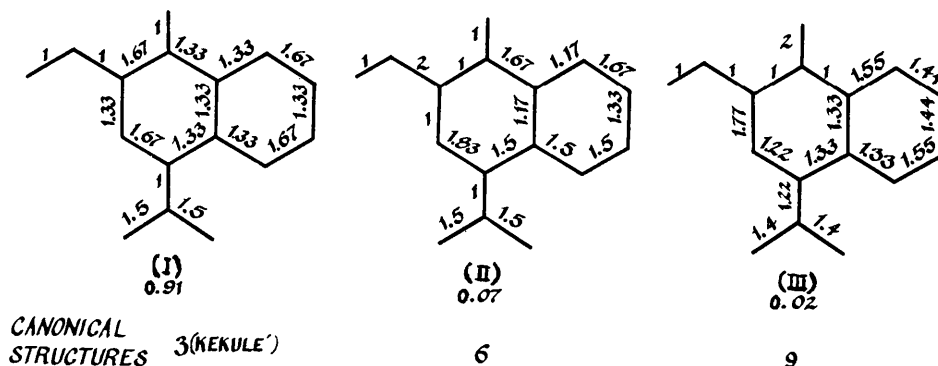
Fig. 3. Molecular packing viewed along the  $b$  axis.







which were derived on the assumption that canonical structures involving independent resonance of the quinoline ring and the substituents contribute 91% to the hybrid structure, while structures with double bonding between the ring and either the acetoxy or methoxy group contribute 2 and 7% respectively. The three classes of resonance structures are represented summarily by the composite structures (I), (II) and (III):



by Dewar & Gleicher (1966) (Table 4). Moreover, the C(8)-N(11) and C(5)-O(16) bond numbers show that there is no major degree of resonance between the nitro or acetoxy group and the quinoline ring. Only a small amount of resonance occurs between the methoxy group and the ring system, as indicated by the C(6)-O(14) bond distance.

The dihedral angles which are listed in Table 5 and the valency angles in Table 6 show the effects of intra-

The bond numbers are the average of those in the equally weighted canonical structures that contribute to the composite structure. The bond lengths of the ring are virtually unaffected by resonance with the substituent groups in that the observed values agree within the uncertainty limits with theoretical distances in the unsubstituted molecule as obtained recently in a self-consistent-field molecular-orbital type calculation

molecular steric repulsions. The exocyclic COC angles in the methoxy and acetoxy groups are significantly larger than tetrahedral due mainly to steric repulsions. In the case of the methoxy group this repulsion is between the methyl group and atoms C(6), C(7) and H(7); in the acetoxy group C(17) and O(18) interact sterically with C(5) of the ring. Similarly, the deviation of the exocyclic angles at C(6) from 120° is due to the steric

Table 3. *Thermal parameters and e.s.d.'s*

Thermal parameters are in the form  $\exp[-2\pi^2(h^2a^*U_{11} + \dots + 2klb^*c^*U_{23})]$ .

	$U_{11}$	$U_{22}$	$U_{33}$	$U_{12}$	$U_{13}$	$U_{23}$
N(1)	0.0465 (0.0019)	0.0360 (0.0020)	0.0359 (0.0018)	-0.0007 (0.0016)	0.0128 (0.0015)	0.0008 (0.0015)
C(2)	0.0549 (0.0019)	0.0475 (0.0017)	0.0438 (0.0022)	-0.0067 (0.0015)	0.0174 (0.0015)	0.0066 (0.0015)
C(3)	0.0562 (0.0016)	0.0329 (0.0015)	0.0585 (0.0018)	-0.0090 (0.0013)	0.0146 (0.0013)	0.0112 (0.0013)
C(4)	0.0439 (0.0013)	0.0260 (0.0014)	0.0501 (0.0015)	-0.0048 (0.0011)	0.0065 (0.0011)	-0.0004 (0.0011)
C(5)	0.0358 (0.0016)	0.0243 (0.0015)	0.0371 (0.0014)	-0.0026 (0.0012)	0.0041 (0.0012)	-0.0041 (0.0011)
C(6)	0.0449 (0.0016)	0.0294 (0.0014)	0.0300 (0.0015)	-0.0007 (0.0012)	0.0044 (0.0012)	-0.0040 (0.0012)
C(7)	0.0497 (0.0014)	0.0225 (0.0014)	0.0335 (0.0014)	-0.0026 (0.0011)	0.0060 (0.0011)	0.0003 (0.0011)
C(8)	0.0382 (0.0013)	0.0242 (0.0014)	0.0333 (0.0014)	-0.0000 (0.0010)	0.0050 (0.0010)	-0.0052 (0.0011)
C(9)	0.0320 (0.0013)	0.0252 (0.0013)	0.0324 (0.0015)	0.0006 (0.0011)	0.0032 (0.0011)	-0.0009 (0.0011)
C(10)	0.0324 (0.0014)	0.0241 (0.0012)	0.0379 (0.0013)	-0.0026 (0.0010)	0.0037 (0.0010)	-0.0017 (0.0010)
N(11)	0.0471 (0.0022)	0.0236 (0.0015)	0.0355 (0.0013)	-0.0011 (0.0014)	0.0100 (0.0013)	-0.0064 (0.0011)
O(12)	0.0968 (0.0015)	0.0434 (0.0013)	0.0374 (0.0019)	-0.0008 (0.0011)	-0.0083 (0.0013)	-0.0105 (0.0012)
O(13)	0.0564 (0.0018)	0.0312 (0.0013)	0.0723 (0.0012)	-0.0134 (0.0012)	0.0072 (0.0011)	-0.0090 (0.0010)
O(14)	0.0797 (0.0022)	0.0359 (0.0019)	0.0323 (0.0018)	-0.0032 (0.0016)	0.0173 (0.0016)	-0.0033 (0.0015)
C(15)	0.0650 (0.0011)	0.0446 (0.0011)	0.0423 (0.0012)	-0.0008 (0.0008)	0.0149 (0.0009)	0.0108 (0.0009)
O(16)	0.0379 (0.0015)	0.0263 (0.0014)	0.0416 (0.0014)	0.0002 (0.0011)	0.0025 (0.0011)	-0.0134 (0.0011)
C(17)	0.0403 (0.0012)	0.0274 (0.0014)	0.0315 (0.0016)	0.0031 (0.0010)	0.0051 (0.0011)	0.0016 (0.0011)
O(18)	0.0400 (0.0022)	0.0414 (0.0023)	0.0587 (0.0019)	-0.0040 (0.0017)	0.0013 (0.0016)	-0.0060 (0.0016)
C(19)	0.0672 (0.0014)	0.0503 (0.0015)	0.0431 (0.0014)	0.0178 (0.0011)	0.0002 (0.0010)	-0.0157 (0.0011)
H(2)	0.0503	0.0523	0.0507	0.0000	0.0000	0.0000
H(3)	0.0503	0.0523	0.0507	0.0000	0.0000	0.0000
H(4)	0.0503	0.0523	0.0507	0.0000	0.0000	0.0000
H(7)	0.0503	0.0523	0.0507	0.0000	0.0000	0.0000
H(24)	0.0503	0.0523	0.0507	0.0000	0.0000	0.0000
H(25)	0.0503	0.0523	0.0507	0.0000	0.0000	0.0000
H(26)	0.0503	0.0523	0.0507	0.0000	0.0000	0.0000
H(27)	0.0503	0.0523	0.0507	0.0000	0.0000	0.0000
H(28)	0.0503	0.0523	0.0507	0.0000	0.0000	0.0000
H(29)	0.0503	0.0523	0.0507	0.0000	0.0000	0.0000

Table 4. *Bond lengths and estimated standard deviations*

$D(\text{obs})$  is the experimental bond length.  $D(\text{adj})$  is the experimental bond length after applying an estimated adjustment for thermal motion.  $D(\text{calc})$  is the bond length that is predicted by the valence bond theory for the corresponding tabulated bond number, B.N.

$D(D-G)$  are theoretical bond lengths in quinoline calculated by an SCF-MO method (Dewar & Gleicher, 1966). The distances and their e.s.d.'s are in Å.

Bond	$D(\text{obs})$	e.s.d.	$D(\text{adj})$	B.N.	$D(\text{calc})$	$D(D-G)$
N(1)—C(2)	1.314	0.005	1.314	1.653	1.298	1.314
C(2)—C(3)	1.408	0.005	1.408	1.335	1.422	1.424
C(3)—C(4)	1.365	0.006	1.365	1.663	1.372	1.374
C(4)—C(10)	1.416	0.004	1.416	1.326	1.424	1.424
C(10)—C(5)	1.422	0.005	1.422	1.350	1.419	1.429
C(5)—C(6)	1.379	0.004	1.379	1.607	1.379	1.373
C(6)—C(7)	1.410	0.004	1.410	1.319	1.425	1.426
C(7)—C(8)	1.371	0.005	1.371	1.670	1.371	1.370
C(8)—C(9)	1.416	0.004	1.416	1.321	1.425	1.436
C(9)—C(10)	1.423	0.004	1.423	1.320	1.425	1.409
C(9)—N(1)	1.370	0.004	1.371	1.345	1.354	1.352
C(8)—N(11)	1.469	0.004	1.472	1.004	1.449	
N(11)—O(12)	1.214	0.004	1.245	1.498	1.260	
N(11)—O(13)	1.229	0.004	1.252	1.498	1.260	
C(6)—O(14)	1.358	0.004	1.373	1.07	1.381	
O(14)—C(15)	1.430	0.005	1.431	1.003	1.432	
C(5)—O(16)	1.386	0.004	1.389	1.02	1.395	
O(16)—C(17)	1.362	0.004	1.365	1.12	1.369	
C(17)—O(18)	1.193	0.004	1.211	1.88	1.219	
C(17)—C(19)	1.492	0.006	1.513	1.00	1.504	
C(15)—H(24)	0.98					
C(15)—H(25)	0.98					
C(15)—H(26)	0.93					
C(19)—H(27)	1.04					
C(19)—H(28)	1.17					
C(19)—H(29)	0.97					
C(2)—H(2)	0.99					
C(3)—H(3)	1.04					
C(4)—H(4)	1.01					
C(7)—H(7)	0.99					

Table 5. *Some least-squares planes of atoms and some dihedral angles*

Plane	Coefficients in $AX+BY+CZ-D=0$ referred to the crystallographic axes ( $X, Y, Z$ in Å)			
	$A$	$B$	$C$	$D$
Quinoline ring	0.9163	-0.1243	0.2595	3.3862
Pyridine moiety	0.9109	-0.1327	0.2702	2.3769
Benzenoid moiety	0.9226	-0.1138	0.2466	3.3508
Nitro group	0.7952	0.3723	-0.5771	-0.6339
Acetoxy group	0.1612	0.6890	-0.7214	-1.2715

	Deviations from the plane (Å)				
	Quinoline	Pyridine	Benzene	Acetoxy	Nitro
N(1)	-0.000	0.002	0.045		
C(2)	0.013	-0.002	0.078		
C(3)	0.015	-0.000	0.082		
C(4)	0.002	0.002	0.051		
C(5)	-0.017	0.019	-0.012	-0.170	
C(6)	0.020	0.072	0.004		
C(7)	0.027	0.080	0.011		-0.035
C(8)	-0.020	0.016	-0.016		
C(9)	-0.018	-0.000	0.008		
C(10)	-0.020	-0.002	0.006		
O(16)				0.001	
C(17)				0.002	
O(18)				-0.004	
C(19)				0.001	

Dihedral angles		Angle
Intersecting planes		79.1°
Acetoxy and benzenoid		59.2
Nitro and benzenoid		1.8
Pyridine and benzenoid		

repulsion between the methoxy group and the ring. The intramolecular contacts involved in these steric interactions are included in Table 7. Certain of the angles in the ring differ significantly from their mean value of  $120^\circ$ . However, the angles C(10)–C(9)–N(1), C(9)–N(1)–C(2) and N(1)–C(2)–C(3) are normal for this heterocyclic system (Pauling, 1960). Repulsion between O(12) and the unshared electron pair on N(1) prevents the nitro group from lying in the plane of the quinoline ring (Fig. 2). The  $59^\circ$  rotation out of the plane produces a normal van der Waals contact of  $2.929 \text{ \AA}$  between N(1) and O(12). While this angle corresponds to an estimated resonance energy of  $1.7 \text{ kcal.mole}^{-1}$  between the nitro group and the ring, the C(8)–N(11) distance indicates negligible resonance

(Trotter, 1960). Presumably packing forces and intramolecular steric requirements govern the magnitude of the out-of-plane rotation, while resonance plays an insignificant role. Repulsions between either O(18) or the C(19)H<sub>3</sub> methyl group and O(14) or H(4) prevent the acetoxy group from lying in the plane of the ring. The rotation of this group by  $79^\circ$  out of the plane decreases the O(18)···C(5) contact to  $2.639 \text{ \AA}$ , about  $0.3 \text{ \AA}$  less than a normal van der Waals distance. Evidently, resonance energy maintains the planarity of the acetoxy group to the extent of forcing this somewhat close intramolecular contact. These resonance effects are evident also in the C(17)–O(18) and C(17)–O(16) bond lengths and numbers.

Equations of some least-squares planes of atoms and the displacements of atoms out of the plane are listed in Table 5. The method of Schomaker, Waser, Marsh & Bergman (1959) was used for computing the equations. The quinoline ring is significantly non-planar although the largest atomic out of plane displacement is only  $0.027 \text{ \AA}$ . While the pyridine moiety is planar to within  $0.002 \text{ \AA}$ , the six membered ring of carbon atoms bearing the three bulky substituent groups is significantly non-planar. Very likely the steric interactions of the substituents induce the deviations from planarity. The intermolecular contacts are at the normal van der Waals distances except for the somewhat close  $2.31 \text{ \AA}$  contact between O(18) of one molecule and H(7) in another molecule related to it by a symmetry center and unit translations in the **b** and **c** directions. The molecular packing is shown in Figs. 2 and 3.

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Table 6. Valency angles and their standard deviations

	Angle	e.s.d.
C(9)–N(1)–C(2)	116.2°	0.3°
N(1)–C(2)–C(3)	124.6	0.4
C(2)–C(3)–C(4)	119.7	0.4
C(3)–C(4)–C(10)	118.5	0.4
C(4)–C(10)–C(9)	117.4	0.3
C(4)–C(10)–C(5)	123.3	0.3
C(9)–C(10)–C(5)	119.2	0.3
C(10)–C(5)–C(6)	121.9	0.3
C(5)–C(6)–C(7)	119.3	0.3
C(6)–C(7)–C(8)	119.1	0.3
C(7)–C(8)–C(9)	123.9	0.3
C(8)–C(9)–C(10)	116.5	0.3
C(8)–C(9)–N(1)	119.9	0.3
C(10)–C(9)–N(1)	123.4	0.3
C(9)–C(8)–N(11)	119.1	0.3
C(7)–C(8)–N(11)	116.9	0.3
C(8)–N(11)–O(13)	117.4	0.3
C(8)–N(11)–O(12)	117.9	0.3
O(12)–N(11)–O(13)	124.6	0.3
C(7)–C(6)–O(14)	124.6	0.3
C(5)–C(6)–O(14)	116.1	0.3
C(6)–O(14)–C(15)	118.1	0.3
C(10)–C(5)–O(16)	118.3	0.3
C(6)–C(5)–O(16)	119.7	0.3
C(5)–O(16)–C(17)	116.7	0.3
O(16)–C(17)–O(18)	122.4	0.4
O(16)–C(17)–C(19)	110.5	0.3
O(18)–C(17)–C(19)	127.0	0.4

Table 7. Some intramolecular distances between non-bonded atoms

Atoms	Distance
C(15)···C(6)	2.392 Å
C(15)···C(7)	2.835
C(15)···H(7)	2.68
H(24)···C(6)	2.71
H(24)···C(7)	2.78
H(24)···H(7)	2.40
H(25)···H(7)	2.40
H(26)···C(6)	2.57
H(26)···C(7)	2.70
O(16)···C(6)	2.39
C(17)···C(5)	2.339
C(17)···O(14)	3.013
C(17)···C(6)	3.068
O(18)···C(5)	2.639
O(18)···C(6)	3.139